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effusion rates for nitrogen (N<sub>2</sub>) and neon (Ne). Rate N / Rate Ne = 0.849 Rate N / Rate Ne =  $\sqrt{\frac{\text{molar mass Ne}}{\text{molar mass N}_2}}$  =  $\sqrt{\frac{20.18}{28.02}}$  = 0.849

2. Calculate the ratio of diffusion rates for carbon monoxide and carbon dioxide. 1.25 Rate CO / Rate CO<sub>2</sub> =  $\sqrt{\frac{\text{molar mass CO}_2}{\text{molar mass CO}}}$

## *States of Matter*

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